

CHEMICAL FORMULA PRACTICE #2

Name the following compounds and *classify* as binary ionic (BI), binary molecular (BM), ternary ionic (TI), or other (more than three symbols).

The highlighted formulas are covered in the binary molecular chapter.

The questions marked with *** are for Chem 1H only.

- 1) potassium nitride
 - 2) potassium nitrate
 - 3) potassium nitrite
 - 4) aluminum hydroxide
 - 5) diphosphorus trioxide
 - 6) strontium cyanide
 - 7) ammonium fluoride
 - 8) lead(II) iodide
 - 9) tin(IV) sulfide
 - 10) dinitrogen nonoxide
 - 11) tetraphosphorus heptoxide
 - 12) iron(III) perchlorate
 - 13) copper(I) hypochlorite
 - 14) zinc permanganate
 - 15) calcium phosphide
 - 16) *** sodium carbonite
 - 17) *** aluminum oxalate
 - 18) *** lead(IV) thiocyanate
 - 19) *** zinc thiosulfate
 - 20) *** lithium hydrogen sulfate
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Name the following and *classify* as binary ionic (BI), binary molecular (BM), and ternary ionic (TI), or other.

The highlighted formulas are covered in the binary molecular chapter.

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|---|---|
| 21) OF ₂ | 31) CaCrO ₄ |
| 22) Ca(NO ₂) ₂ | 32) Mg(C ₂ H ₃ O ₂) ₂ |
| 23) K ₂ Cr ₂ O ₇ | 33) Cr(HCO ₃) ₃ |
| 24) SCl ₆ | 34) SO ₃ |
| 25) Na ₃ PO ₃ | 35) Na ₂ Se |
| 26) As ₂ Br ₅ | 36) *** MnS ₂ O ₃ |
| 27) NO ₂ | 37) *** Ga ₂ (C ₂ O ₄) ₃ |
| 28) SnO | 38) *** CdCO ₂ |
| 29) (NH ₄) ₃ P | 39) *** Sc(BrO ₃) ₂ |
| 30) Al ₂ (SO ₃) ₃ | 40) *** FeAsO ₄ |